



K_c & ITS UNITS

Equilibrium	K_c	Units
$N_2 + 3H_2 \rightleftharpoons 2NH_3$	$K_c = \frac{[NH_3]^2}{[N_2][H_2]^3}$	$\frac{(mol\ dm^{-3})^2}{(mol\ dm^{-3})^4} = \frac{1}{(mol\ dm^{-3})^2} = \frac{1}{mol^2\ dm^{-6}} = mol^{-2}\ dm^6$
$PCl_3 + Cl_2 \rightleftharpoons PCl_5$		
$H_2 + I_2 \rightleftharpoons 2HI$		
$2SO_3 \rightleftharpoons 2SO_2 + O_2$		
$CH_4 + H_2O \rightleftharpoons 3H_2 + CO$		
$2NO_2 \rightleftharpoons N_2O_4$		